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Professor Paul H. Emmett

Chemistry Department

Portland State University

P.O. Box 751 Portland,

Oregon 97207 June 17, 1977

the fact that nitrogen at —

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195°C is about 5% imperfect at atmospheric pressure, the correction being a linear function of the pressure. Instead of a curve with kinks a smooth S-shaped isotherm was obtained with nitrogen at  $-195^{\circ}\text{C}$ .

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Paul Hugh Emmett was an American chemist best known for his pioneering work in the field of catalysis and



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for his work on the  
Manhattan Project during  
World War II. He spearheaded  
the research to separate  
isotopes of uranium and to  
develop a corrosive uranium  
gas. Emmett also made  
significant contributions to

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BET Theory which explains the relationship between surface area and gas adsorption. He served on the faculty of Johns Hopkins University for 23 years throughout his scientific career.

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~~Paul Hugh Emmett — Wikipedia~~

S. Brunauer, P. H. Emmett  
and E. Teller, "Adsorption  
of Gases in Multimolecular  
Layers," Journal of the  
American Chemical Society,  
Vol. 60, No. 2, 1938, pp.

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~~S. Brunauer, P. H. Emmett  
and E. Teller, "Adsorption  
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The gas adsorption technique  
is based on the Brunauer-  
Emmett-Teller (BET) model

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included in the determination of the amount of gas that is adsorbed on the surface of the sample at low temperatures [13]. The specific surface area ( $m^2/g$ ) defined by this method comprises the outward as

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well as the internal (open pores) surface area.

However, the surface area of closed pores cannot be determined because of its inaccessibility for adsorbing gas molecules.

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theory of multimolecular  
adsorptionis extended to  
mixtures of gases. No  
satisfactory experimental  
data are available at the  
present time with which to  
test the theory.

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Emmett PH, Brunauer S, Love  
KS. The measurement of  
surface areas of soils and  
soil colloids by the use of

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low temperature van der  
waals adsorption isotherms  
Soil Science. 45: 57-65. 1:  
1937: Emmett PH, Brunauer S.  
Accumulation of alkali  
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iron synthetic ammonia  
catalysts Journal of the

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American Chemical Society.  
59: 310-315. 1: 1937

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~~Publications~~

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papers showed that various experimental adsorption isotherms could be fitted by equation (4a). However, they could not evaluate  $K_1$  because in all cases, except one, the surface, and therefore  $v_m$ , was not known.

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theory aims to explain the physical adsorption of gas molecules on a solid surface and serves as the basis for an important analysis technique for the measurement of the specific surface area of materials.

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The observations are very often referred to as physical adsorption or physisorption. In 1938, Stephen Brunauer, Paul Hugh Emmett, and Edward Teller published the first article about the BET theory in the

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Journal of the American  
Chemical Society. The BET  
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Stephen Brunauer, P. H.  
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Chemical Society 1938, . . .  
Herbert S. Harned; Journal  
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Brunauer, S., Emmett, P.H.

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(BET) equation (Brunauer et al., 1938) overcomes the limitation to a monolayer of sorbate and currently is a commonly used model to describe the multilayer adsorption of gas molecules on a sorbent (Brunauer et



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