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Particulate nature of

matter - Part-1 -

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States of Matter

Chemical and Physical

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Liquid Gas *10 Amazing
Experiments with Water*

Pure Substances and
Mixtures!

(Classification of
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Understanding Atoms,
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formation,
chlorofluorocarbons,
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and Matter Chemistry
Central Science

Chapter Summaries – Chemistry Matter and Change

Chapter 1 Section 2

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MATTER and Its

Properties Matter –

Anything that has

MASS and VOLUME.

ATOMS and

MOLECULES the basic building blocks of

matter. **ATOM--** the

smallest unit of an

element that still

maintains the properties

that element.

ELEMENT– a pure

substance made up of

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only one type of atom.

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Section 1

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hanges Matter—Propertie
s and Changes Solutions

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Chapter 3 35 Section 3.1

Properties of Matter

pages 70–75 Problem-

Solving Lab 1. Explain

why the flow of a

compressed gas must be

controlled for practical

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and safe use.

**Matter—Properties and
Changes**

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Periodic Law Section

6.1 Development of the Modern Periodic Table

1. octaves
2. eight
3. nine
4. accepted
5. Dmitri Mendeleev
6. atomic mass
7. elements
8. atomic number
9. Henry Moseley
10. protons
11. periodic law
12. properties
13. 14 ...

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Chapter 5 25 Electrons

in Atoms Reviewing

Vocabulary Match the

definition in Column A

with the term in Column

B. Column A Column B

1. The set of frequencies

of the electromagnetic

waves emitted by the

atoms of an element 2.

The minimum amount

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of energy that can be
lost or gained by ...
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of Matter - 2.1 Lesson

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extensive property.

intensive property.

extensive properties

(examples) intensive

property (examples) a

property that depends on

the amount of matter in

a sample. a property that

depends on the type of

matter in a sample not....

mass, volume, length,

weight.

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number of moles of

known substance C H C

H C H O O O CO CO

CO H O H O H O

number of moles of

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unknown substance O

CO H O C H CO H O C

H O H O C H O CO use

mole ratio 5 mol O 2 1

mol C 3 H 8 3 mol CO 2

1 mol C 3 H 8 4 mol H

2 O 1 mol C 3 H 8 1

mol C 3 H 8 5 mol O 2

MATH SKILLS

TRANSPARENCY

MASTER 16

A “chemical change” is
a change that produces

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matter with a different composition than the original matter.

Chemical Change A change in which one or more substances are converted into different substances. Heat and light are often evidence of a chemical change.

Properties of Compounds Quite different properties than their component

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